

Thermodynamic analysis of glycerol conversion to synthesis gas

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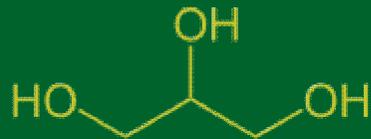
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Why glycerol? Why synthesis gas?

Glycerol



- unique structure and properties
- bioavailability
- renewability

- biodiesel production**
(main side-product: ~ 10 wt. %)
- numerous possibilities for its conversion to:
 - acrolein and acetol
 - diols (1,2- and 1,3-propanediol, ethylene glycol)
 - glycerol carbonate
 - **synthesis gas**

Synthesis gas



- ❖ many applications in chemical synthesis
 - methanol production
 - hydrogen production
 - oxosynthesis
 - Fischer-Tropsch process



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Main reagents of glycerol conversion (with or without CO₂):

Glycerol, carbon monoxide, carbon dioxide, hydrogen, water, methane

Table 1. Selected stoichiometric models.

1 st model	$\text{C}_3\text{H}_8\text{O}_3 + 3\text{H}_2\text{O} = 3\text{CO}_2 + 7\text{H}_2 \text{ (1a)}$ $\text{CO}_2 + \text{H}_2 = \text{CO} + \text{H}_2\text{O} \text{ (1b)}$ $\text{CO}_2 + 4\text{H}_2 = \text{CH}_4 + 2\text{H}_2\text{O} \text{ (1c)}$
2 nd model	$\text{C}_3\text{H}_8\text{O}_3 + \text{CO}_2 = 4\text{CO} + 3\text{H}_2 + \text{H}_2\text{O} \text{ (2a)}$ $\text{CO}_2 + \text{H}_2 = \text{CO} + \text{H}_2\text{O} \text{ (2b)}$ $\text{CO} + 3\text{H}_2 = \text{CH}_4 + \text{H}_2\text{O} \text{ (2c)}$
3 rd model	$\text{C}_3\text{H}_8\text{O}_3 = 3\text{CO} + 4\text{H}_2 \text{ (3a)}$ $\text{CO} + 3\text{H}_2 = \text{CH}_4 + \text{H}_2\text{O} \text{ (3b)}$ $\text{CO} + \text{H}_2\text{O} = \text{CO}_2 + \text{H}_2 \text{ (3c)}$

Glycerol conversion with steam
(steam reforming)

Glycerol conversion with carbon dioxide
(dry reforming)

Glycerol conversion without any co-substrate

Conclusions



It was found that the contribution of synthesis gas in the equilibrium mixture increases with the increase temperature.



In the steam reforming of glycerol synthesis gas rich with hydrogen is produced, while in the dry reforming the carbon monoxide predominates in the equilibrium mixture.

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